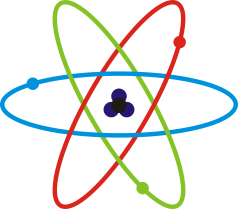
[](http://www.google.com/url?sa=i&source=images&cd=&cad=rja&docid=qqRhthjJCxPZyM&tbnid=cxjvQsoPiDzfaM:&ved=0CAgQjRwwAA&url=https%3A%2F%2Fen.m.wikipedia.org%2Fwiki%2FFile%3ASchematicky_atom.png&ei=6FldUvmSE7fF4APlwoCgDA&psig=AFQjCNHvlr2ek82FB47y25gL2XuyncIAMg&ust=1381935976352072)**Unit 3: Atomic Structure**

**Each element on the periodic table represents a different type of atom, let’s take a look at what an atom is made of:**

Structure of the Atom

Protons:

Neutrons:

Electrons:

How to determine the number of protons, neutrons, and electrons:

Examples:

Al Ar Mg+2 F-1

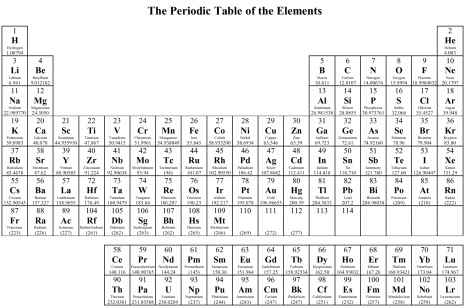
P P P P

N N N N

E E E E

**Ion:**

**Isotope:**



**Important vocab concepts that we will be using throughout the year**:

Metals-

Alkali metals Alkaline earth metals- Transition elements-

Nonmetals-

Halogens- Noble gases-

Metalloids/ semi-metals-

Bonding:

Who?

Why?

How?

Type of Bonds

|  |  |
| --- | --- |
| Covalent |  |
| Nonpolar covalent |  |
| Polar covalent |  |
| Ionic |  |

Properties

|  |  |
| --- | --- |
| Ionic | Covalent |
|  |  |